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Unit 4 Atoms and the Periodic Table  
Study Guide 1. The smallest

“ uncuttable ” particle of an element is  
the \_\_\_\_\_. 2. Atoms of the same element

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Answers have exactly the same \_\_\_\_\_. Atoms of different elements have different \_\_\_\_\_. 3.

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~~Answer~~ & the Periodic Table; PhET Simulation:  
Build an Atom; Exploring the Periodic  
Table; Extra Practice - Using the Periodic  
Table; ChemThink: Atomic Structure;  
Video: How does an Atom look like? Ions  
& Ion Notation (Atom with a Charge)

~~3. Atoms & the Periodic Table - Mrs.~~

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## Summary

The periodic table helps predict some properties of the elements compared to each other. Atom size decreases as you move from left to right across the table and increases as you move down a column. The energy required to remove an electron from an atom increases as you

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~~Answer~~ move from left to right and decreases as you move down a column.

~~Periodic Table Study Guide Introduction  
& History~~

Honors Study Guide- Atoms and the  
Periodic Table of Elements 1. Define  
Element: An element is a chemical

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**Answer** substance that is in the form of an atom 2.  
Define Atom: An atom is small sub-atomic particles held together by fundamental forces 3. Describe changes in the atomic model over time and why those changes were necessitated by experimental evidence. Explain how each discovery led to the modern ...

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Study Guide – Atoms & the Periodic  
Table Name \_\_\_\_\_ The Scientists A.

Thomson ' s study of cathode rays  
revealed that atoms contained negatively  
charged particles called \_\_\_\_\_. B.



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Rutherford ' s gold-foil experiment revealed that an atom ' s positive charge is concentrated in the atom ' s

~~Study Guide Atoms & the Periodic Table~~  
Name

The position of elements on the periodic table is directly related to their electron

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~~Answer~~ Describe the relationship between group number and number of valence electrons. the group number is how many valence electrons of that element are in each atom. Describe the role of a coefficient in a chemical equation.

~~Atoms and the Periodic Table Study~~

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~~Answer Flashcards | Quizlet~~

Study Guide: Atoms & the Periodic Table

Subpages (7): Lesson 1 Lesson 1: Parts of  
the Atom Lesson 2: History of the Atom

Lesson 3: Ions & Isotopes Lesson 4:  
Electron Location/Configuration Lesson...

~~Unit 4: Atoms & the Periodic Table~~

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~~SMHS Physical Science~~

KS3 Chemistry Atoms, elements and the periodic table learning resources for adults, children, parents and teachers.

~~Atoms, elements and the periodic table -  
KS3 Chemistry ...~~

The study guide for this unit is provided

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here: Study Guide - Atoms and Elements

4) Daily Warm Ups - Daily warm ups, "do it nows" or "bell activities" serve the purpose of engaging students in the daily learning, focusing students for scientific study and to practice skills or learn concepts.

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~~Study Guide~~ ~~Atoms and Elements~~  
~~Better Lesson~~

Mass number. sum of the number of protons and neutrons; number of protons + number of neutrons = mass number, number of neutrons = mass number - number of protons (atomic number), number of protons = number of electrons

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~~Answer~~ IF positive and negative charges cancel,  
the atom Charge= 0 (Stable state, non  
ionic state) Isotopes.

~~Unit 1: atoms & the periodic Table; Atoms  
Flashcards | Quizlet~~

Atoms and the Periodic Table Study  
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Matter Unit Information

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6 Test The Periodic Table.

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~~Answer~~

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Electron affinity increases from left to right

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**Answer** across the periodic table. This is caused by the decrease in atomic radius. As we already explained, moving from left to right across a period, atoms become smaller and smaller as the atomic number increases. This causes the electron to move closer to the nucleus, thus increasing the electron affinity.

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Atoms and the Periodic Table Study  
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and password.



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## ~~Quia Study Guide for Atoms and the Periodic Table~~

Each row in the periodic table is called a period and the period number tells how many energy levels atoms in that period have. As a result, as you go down a group in the periodic table the sizes...

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~~Describe atomic size. How does it relate to the periodic ...~~

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